

## Supporting Information

### Figure Legends

#### Figure S1.

FT infra red spectrum of crude acids showing two absorptions attributed to carbonyl stretches of acids ( $1708\text{ cm}^{-1}$ ) and esters ( $1726\text{ cm}^{-1}$ )

#### Figure S2.

Electron ionization mass spectrum of the putative  $\text{C}_{80}$  hexacyclic triacid mono'octyl'ester ('the unknown').

#### Figure S3.

a. Positive ion Orbitrap electrospray mass spectrum of the trimethyl ester of the putative  $\text{C}_{80}$  hexacyclic triacid mono'octyl'ester ('the unknown'). A formula of  $\text{C}_{91}\text{H}_{164}\text{O}_8\text{NH}_4$  is supported by the mass of the ammoniated molecular ion ( $1403.2749$  observed;  $1403.2764$  required).

b. Negative ion Orbitrap electrospray mass spectrum of the putative  $\text{C}_{80}$  hexacyclic triacid mono'octyl'ester ('the unknown'). A formula of  $\text{C}_{88}\text{H}_{157}\text{O}_8$  is supported by the mass of the deprotonated molecular ion ( $1342.1857$  observed;  $1342.1884$  required).

#### Figure S4.

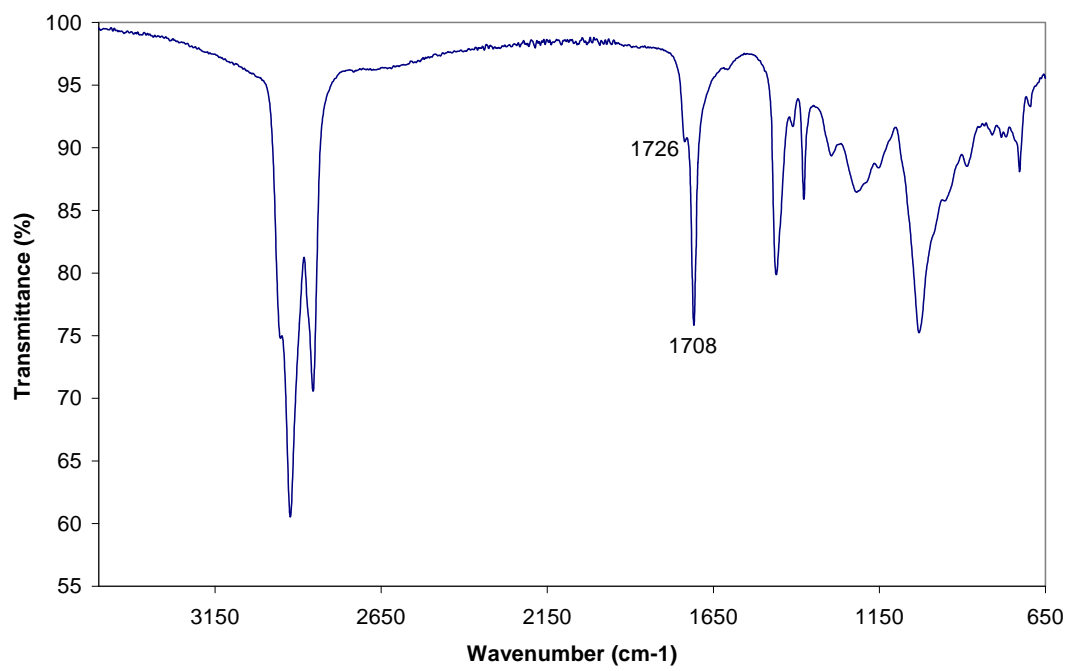
a. Electron ionization mass spectrum and GC retention time of the alcohol (TMS ether) obtained by saponification of the putative  $\text{C}_{80}$  hexacyclic triacid mono'octyl'ester ('the unknown'), followed by examination by GC-MS.

b. Electron ionization mass spectrum and GC retention time of authentic 2-ethylhexanol (TMS ether).

c. Electron ionization mass spectrum and GC retention time of *n*-octanol (TMS ether).



Figure S1.



**Figure S2.**

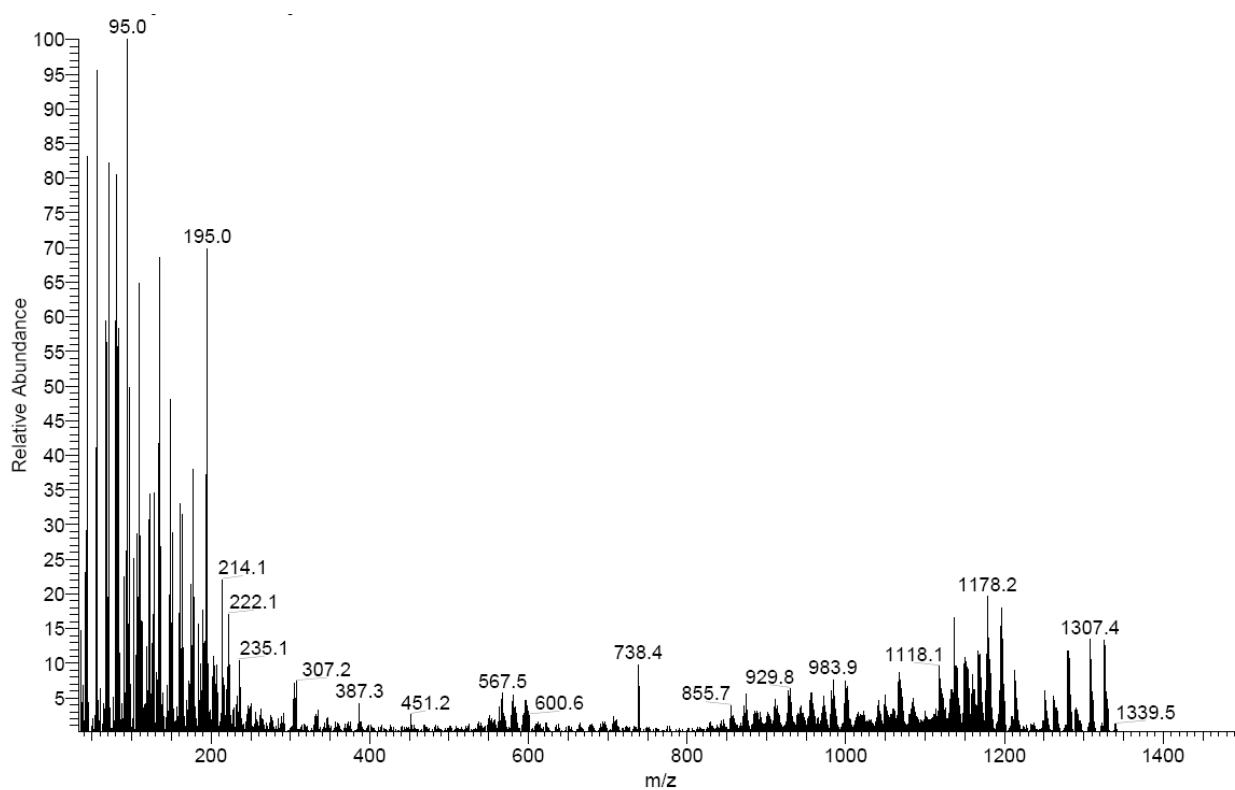
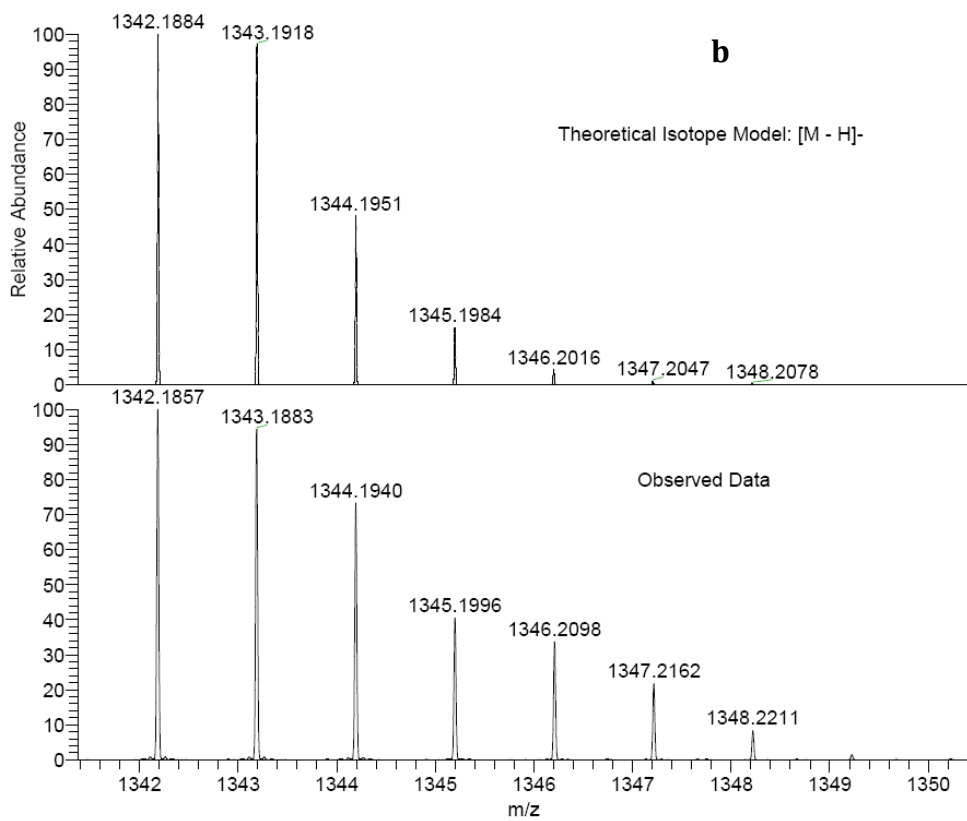
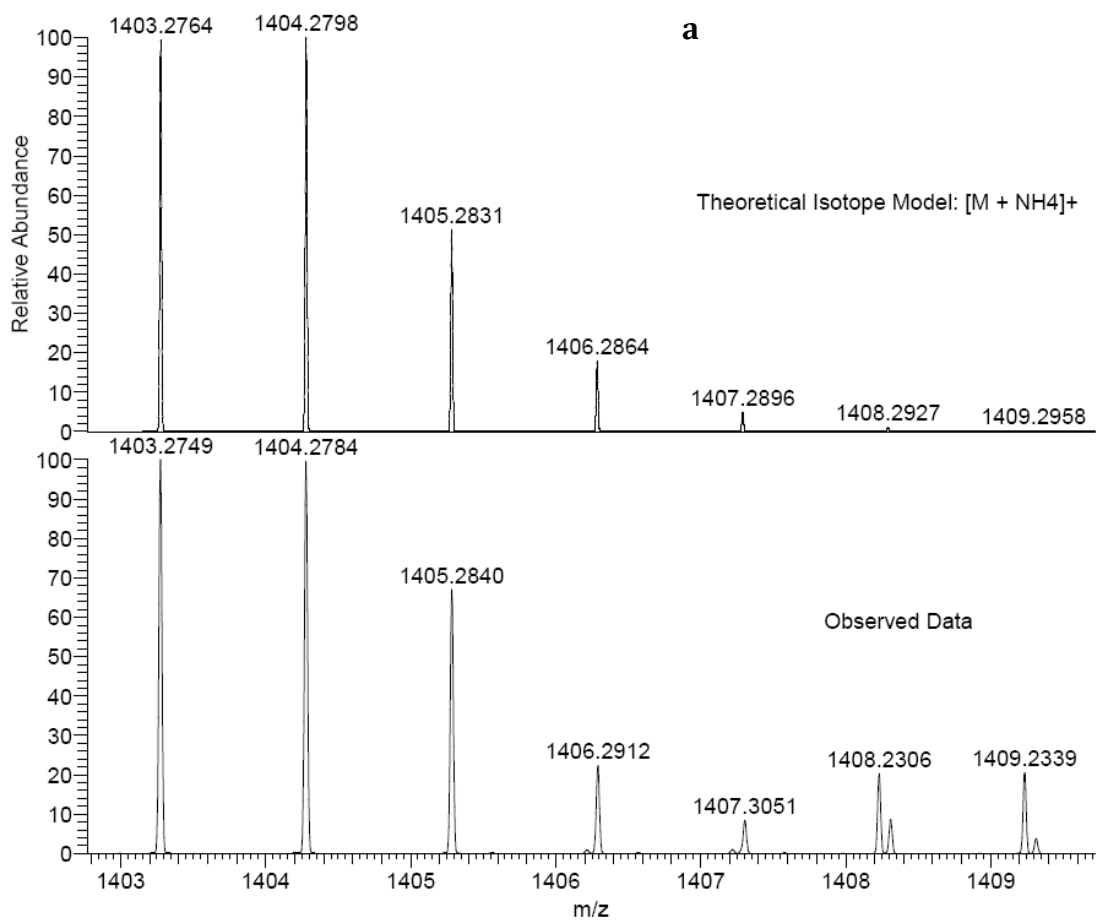
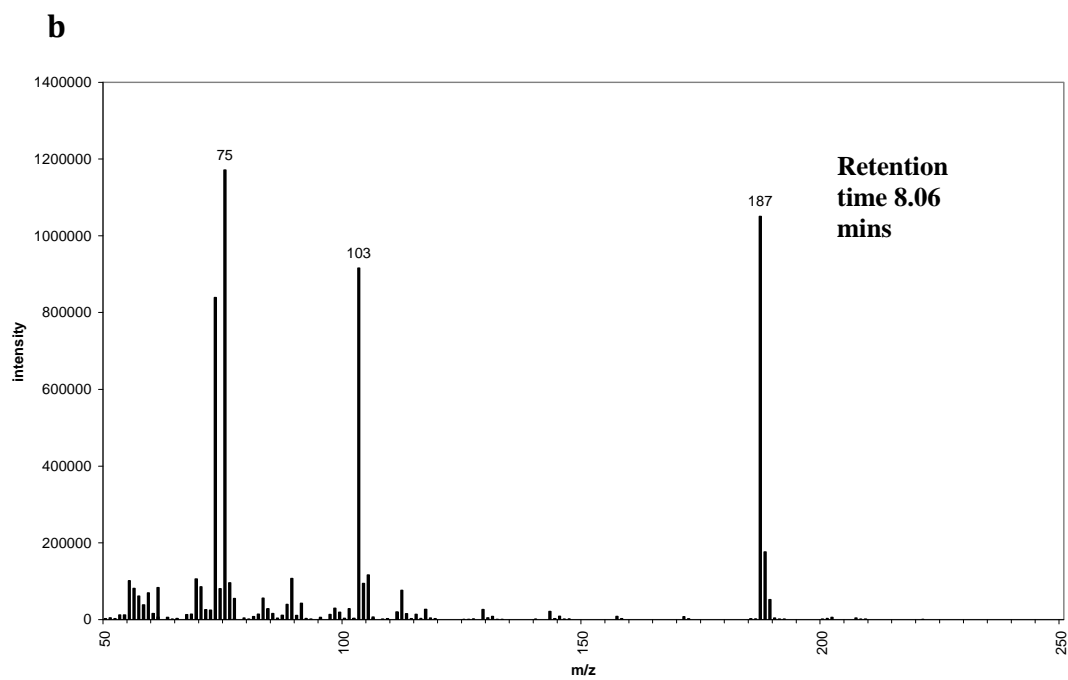
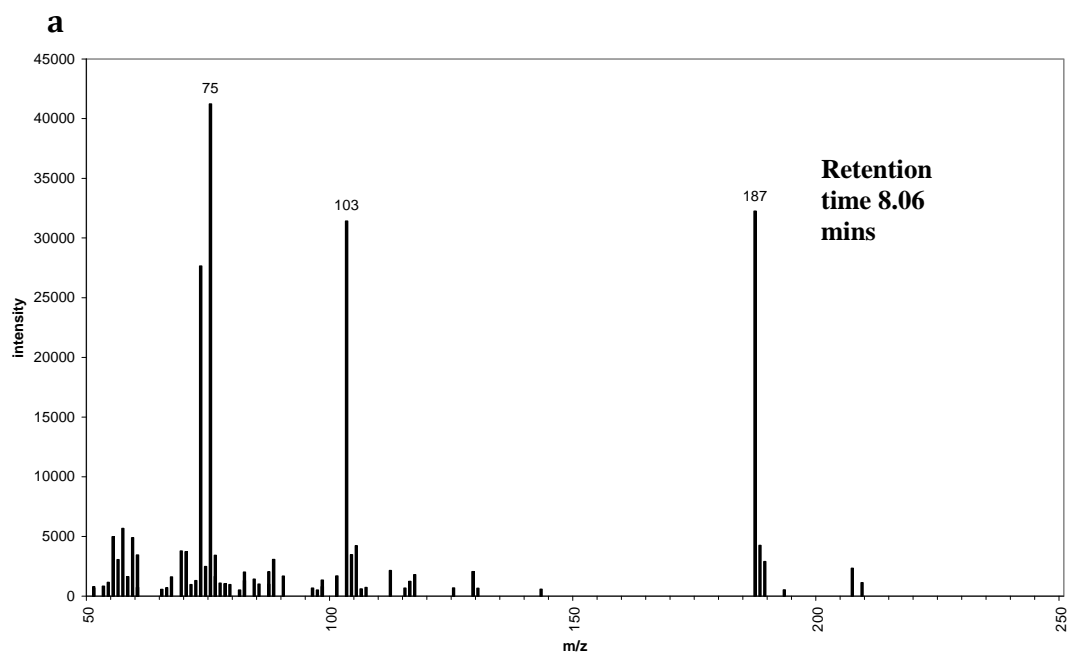


Figure S3.



**Figure S4.**



**C**

