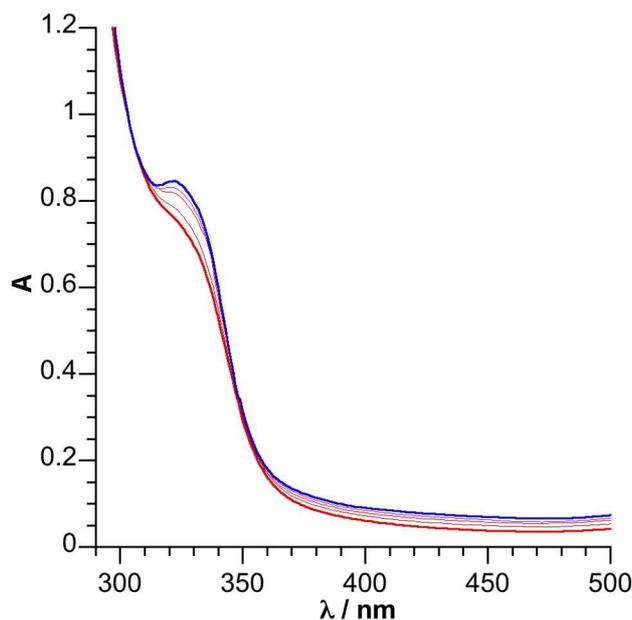


## Photochemical activation of an azido manganese-monosubstituted Keggin polyoxometalate: on the road to a Mn(V)-nitrido derivative.

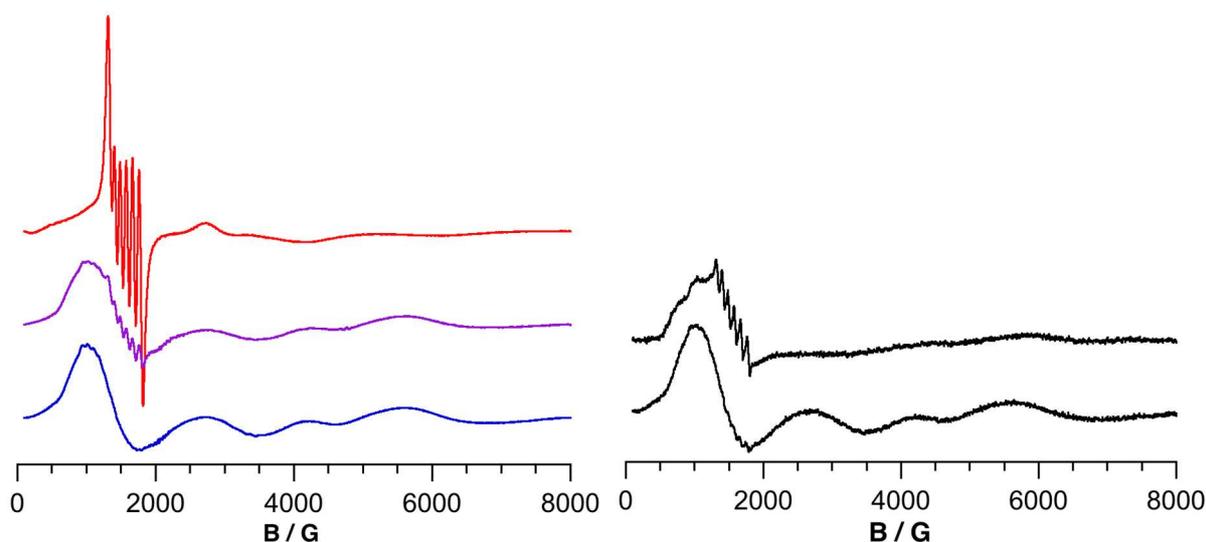
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### Supporting Information

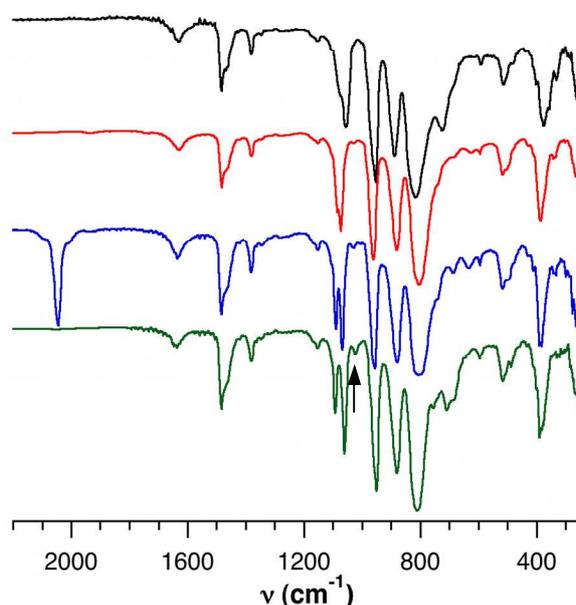


Effect of temperature on the absorbance spectral profile for **TBA-2(N<sub>3</sub>)** in butyronitrile. Spectra were recorded from room temperature (red curve) to 140 K (blue curve).

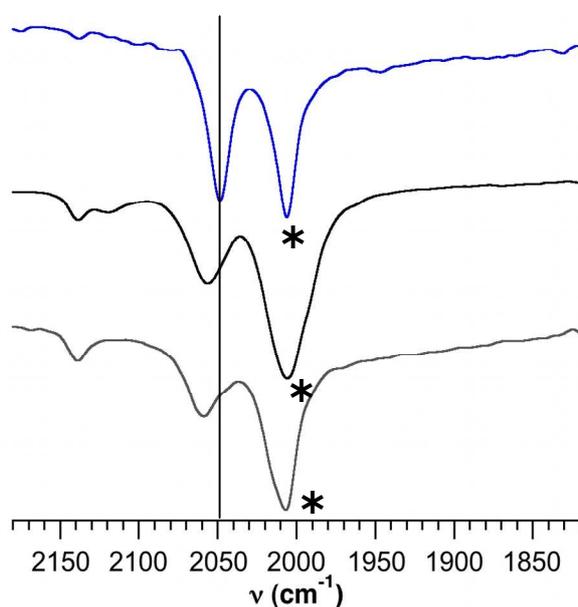


Left: Evolution of the EPR spectrum (X band) of **TBA-1(OH<sub>2</sub>)** (red) upon addition of increasing amounts of **TBaN<sub>3</sub>** in CH<sub>3</sub>CN (77 K). The blue plot corresponds to the signal of **TBA-1(N<sub>3</sub>)** while the purple one corresponds to a mixture of **TBA-1(OH<sub>2</sub>)** and **TBA-1(N<sub>3</sub>)**.

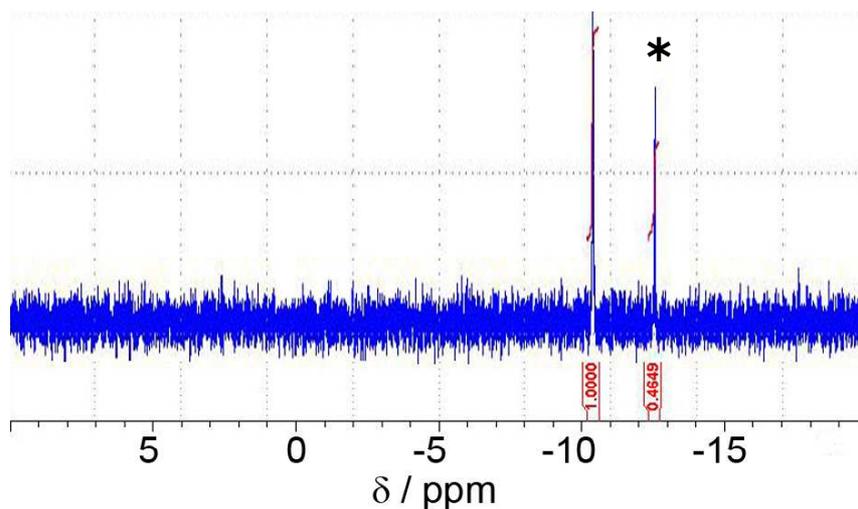
Right up: EPR spectrum (X band) of a photolysed solution (room temperature,  $\lambda_{\text{exc}} = 335 \text{ nm}$ ) of **TBA-2(N<sub>3</sub>)** (0.5 mM, 2 eq. of TBAN<sub>3</sub> per Mn<sup>III</sup>) in CH<sub>3</sub>CN. Right down: same sample after addition of an excess of TBAN<sub>3</sub>.



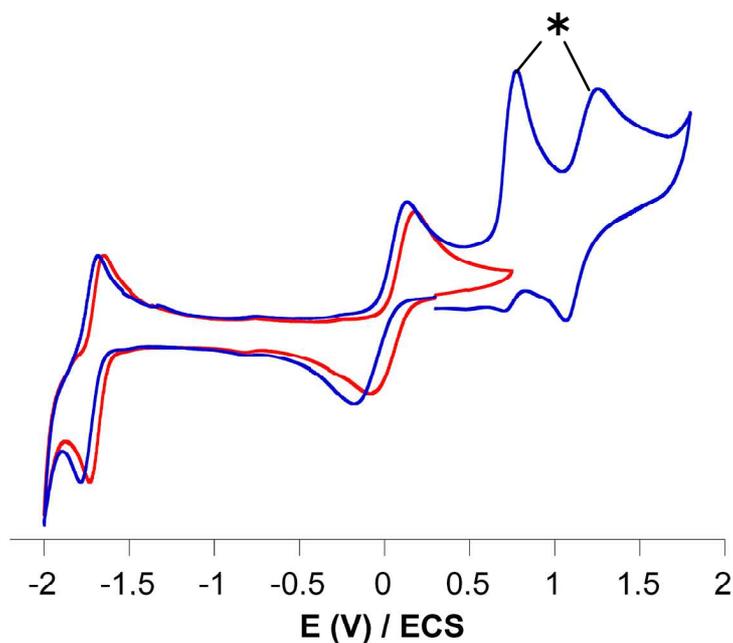
IR spectra of **TBA-1(OH<sub>2</sub>)** (black), **TBA-2(OH<sub>2</sub>)** (red), **TBA-2(N<sub>3</sub>)** (blue) and **TBA-3** in KBr (the arrow indicates the very weak band at 1023 cm<sup>-1</sup> that was assigned to the Mn-N stretch).



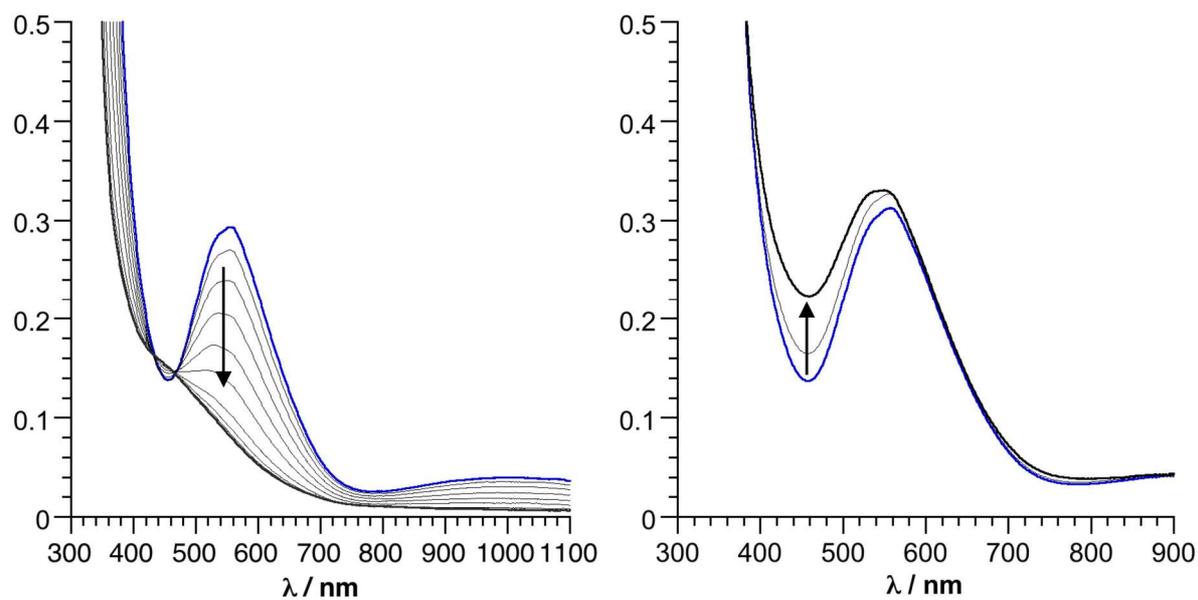
1800-2200 cm<sup>-1</sup> region of the IR spectra of solutions of **TBA-2(N<sub>3</sub>)** (0.5 mM, 2 eq. of TBAN<sub>3</sub> per Mn<sup>III</sup>) in MeCN (blue), **TBA-1(N<sub>3</sub>)** (1 mM, large excess of TBAN<sub>3</sub>) in MeCN (black) and photolysed solution (room temperature,  $\lambda_{\text{exc}} = 335 \text{ nm}$ ) of **TBA-2(N<sub>3</sub>)** (0.5 mM, 5 eq. of TBAN<sub>3</sub> per Mn<sup>III</sup>) in CH<sub>3</sub>CN (grey). The peaks marked by \* corresponds to the free N<sub>3</sub><sup>-</sup> present in solution.



$^{31}\text{P}$  NMR ( $\text{CD}_3\text{CN}/\text{CH}_3\text{CN}$ , 121.5 MHz) spectrum of a solution of **TBA-3** obtained by photolysis at  $-30^\circ\text{C}$  with  $\lambda_{\text{exc}} = 313 \text{ nm}$  of **TBA-2(N<sub>3</sub>)** (1.0 mM, 4 eq. of  $\text{TBAN}_3$  per  $\text{Mn}^{\text{III}}$ ). The peak marked by \* corresponds to  $\text{TBA}_3[\text{PW}_{12}\text{O}_{40}]$  (0.5 mM) used as internal reference.



Cyclic voltammograms of **TBA-2(OH<sub>2</sub>)** (1 mM, red) and **TBA-2(N<sub>3</sub>)** (1.0 mM, 2 eq. of  $\text{TBAN}_3$  per  $\text{Mn}^{\text{III}}$ , blue) in  $\text{CH}_3\text{CN}$  with  $10^{-1} \text{ M}$   $\text{TBABF}_4$  as supporting electrolyte versus SCE (scan rate  $100 \text{ mV}\cdot\text{s}^{-1}$ ). The waves marked by \* correspond to the oxidation of free  $\text{N}_3^-$ .



Left: spectral changes observed during the photolysis of **TBA-2(N<sub>3</sub>)** (1 mM) in CH<sub>3</sub>CN solution with  $\lambda_{\text{exc}} = 365$  nm at room temperature. The photolysis was complete within 1 hour. Right: same experiment carried out at -30°C. After 1 hour of photolysis, (black bold curve) the reaction was still far from completion.